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## Behavior of Real gases.

Real gases attain ideal gas behavior at high temperatures and low pressures because at high pressures and low temperatures, the molecules of gases are very close to each other. Molecular interactions begin to operate. Repulsive forces also become important. Repulsive interactions are short-range interactions and are important when molecules are almost in contact and the assumption for an ideal gas is that there are negligible attractive force between gas molecules. Real gases show ideal behavior when the temperature and pressure conditions are such

that intermolecular forces are practically negligible. It is possible only at very low pressure and very high temperature.